## Balancing Chemical Equations - Practice Part 2

7)  $\underline{\hspace{1cm}}$  Na +  $\underline{\hspace{1cm}}$  O<sub>2</sub>  $\rightarrow$   $\underline{\hspace{1cm}}$  Na<sub>2</sub>O

1)  $\underline{\hspace{1cm}}$  K<sub>3</sub>PO<sub>4</sub> +  $\underline{\hspace{1cm}}$  HCl  $\rightarrow$   $\underline{\hspace{1cm}}$  KCl +  $\underline{\hspace{1cm}}$  H<sub>3</sub>PO<sub>4</sub>

2)  $\underline{\hspace{1cm}}$  MgF<sub>2</sub> +  $\underline{\hspace{1cm}}$  Li<sub>2</sub>CO<sub>3</sub>  $\rightarrow$   $\underline{\hspace{1cm}}$  MgCO<sub>3</sub> +  $\underline{\hspace{1cm}}$  LiF

3)  $P_4 + Q_2 \rightarrow P_2O_3$ 

4) \_\_\_\_  $CF_4 +$  \_\_\_  $Br_2 \rightarrow$  \_\_\_  $CBr_4 +$  \_\_\_  $F_2$ 

5)  $\underline{\qquad}$  GaF<sub>3</sub> +  $\underline{\qquad}$  Cs  $\rightarrow$   $\underline{\qquad}$  CsF +  $\underline{\qquad}$  Ga

6)  $\underline{\hspace{1cm}}$  BaS +  $\underline{\hspace{1cm}}$  PtF<sub>2</sub>  $\Rightarrow$   $\underline{\hspace{1cm}}$  BaF<sub>2</sub> +  $\underline{\hspace{1cm}}$  PtS

8)  $NaF + Br_2 \rightarrow NaBr + F_2$ 

9) \_\_\_\_ S + \_\_\_  $O_2 \rightarrow$  \_\_\_\_  $SO_3$ 

10) \_\_\_\_  $CH_4 +$ \_\_\_  $O_2 \rightarrow$ \_\_\_  $CO_2 +$ \_\_\_  $H_2O$ 

- 11) \_\_\_\_ Na + \_\_\_ HCl  $\rightarrow$  \_\_\_ H<sub>2</sub> + \_\_\_ NaCl
- 12) How many total molecules of product were made in #8?\_\_\_\_\_
- 13) How many total molecules of reactants were used in #5?\_\_\_\_\_
- 14) How many reactant types are used in #11?\_\_\_\_\_
- 15) How many types of products are made in #9? \_\_\_\_\_
- 16) 4 molecules of a substance contains 3 atoms of sodium, 1 atom of phosphorus, and 4 atoms of oxygen. What is the correct chemical formula for this substance?